MODULE HANDBOOK

Module Name	**Practicum of Inorganic Chemistry**
Module Level	Bachelor
Abbreviation, if applicable	3074211051
Sub-heading, if applicable	-
Course included in the module, if applicable	-
Semester/term	6th/ third year
Module coordinator(s)	Kusumawati Dwiningsih, S.Pd., M.Pd.
Lecturer(s)	**Dr. Achmad Lutfi, M.Pd.; Dr. Amaria, M.Si., Prof. dr. Sari Edi C., M.Si, Dr. Muchlis, M.Pd.; Dina Kartika M., S.Si., M.Sc, Kusumawati D., S.Pd.,M.Pd.; Rusly Hidayah, S.Si., M.Pd.**
Language	Bahasa Indonesia
Classification within the curriculum	Compulsory
Teaching format/class hours per week during the semester	1 hours lectures (150 min / hour)
Workload	1 CU for bachelor degree equals to 3 workhours per week or 170 minutes (50' face to face learning, 60' structured learning, and 60' independent learning). In one semester, courses are conducted in 14 weeks (excluding mid and end-term exam). Thus, 1 CU equals to 39.67 workhours per semester. One CU equals to 1.59 ECTS.
Credit point	3 SCU (3x1.59 =4.77 ECTS)
Requirement	General Chemistry II
Learning Outcomes	**General Competence (knowledge)**: Students can understand the physical-chemical properties, preparation of main group elements (alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen) and transition elements **Specific Competence** : At the end of the lecture, students can study physical-chemical properties, preparation of its compounds in laboratory scale of alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen, study preparation of cis trans metal complexes, metal salt complexes, study the strength of ligand fields in metal complexes, and study reactions in metal complexes

Content	Course materials discuss physical-chemical properties, preparation of its compounds in laboratory scale of alkalis, alkaline earth, boron family, carbon family, nitrogen family, oxygen family, halogen and hydrogen, study preparation of cis trans metal complexes, metal salt complexes, study the strength of ligand fields in metal complexes, and study reactions in metal complexes.
Study/exam achievements	Students are considered to be competent and pass if at least get 55 Final score is calculated as follows: 25% experiment + 20% Task + 20% sub summative test + 35% final exam Table index of graduation A = 4 (85 ≤-≥ 100) A- = 3,75 (80 ≤-< 85) B+ = 3,5 (75 ≤- < 80) B = 3 (70 ≤-< 75) B- = 2,75 (65 ≤-<75) C+ = 2,5 (60 ≤-<65) C = 2 (55 ≤-<60) D = 1 (40 ≤-<55) E = 0 (0 ≤-<40)
Media:	Computer, LCD, White board, Chemical Equipment
Learning Methods	Individual assignment, group assignment, discussion, and presentation
Literature:	1. Lee, J.D. 1991. *ConciseInorganic Chemistry*. Four Edition. London: Chapman & Hall. 2. Madan, R.D. 1997. *Modern Inoragnic Chemistry*. New Delhi: S. Chand and Company LDT. 3. Manku, G.S. 1980.*Inorganic Chemistry.* India: Tata Mc Graw Hill Book Co. 4. Sugiarto, B. dkk. 1997. *Kimia Anorganik*. Surabaya: Unipress IKIP Surabaya