MODULE HANDBOOK

Module Name	Coordination Chemistry
Module level	Bachelor
Abbreviation, if applicable	3074212034
Sub-heading, if applicable	-
Course included in the module, if applicable	-
Semester/term	4th semester/Second Year
Module coordinator(s)	Dr. Amaria, M.Si.
Lecturer(s)	Prof. Dr. Sari Edy Cahyaningrum. M.Si., Dr. Amaria, M.Si. Dina Kartika Maharani, S.Si., M.Sc.
Language	Indonesian
Classification within the curriculum	Compulsory Course
Teaching format/class hours per week during the semester:	2 hours lecturers (50 min per hours)
Workload:	1 CU for bachelor degree equals to 3 workhours per week or 170 minutes (50' face to face learning, 60' structured learning, and 60' independent learning). In one semester, courses are conducted in 14 weeks (excluding mid and end-term exam). Thus, 1 CU equals to 39.67 workhours per semester. One CU equals to 1.59 ECTS.
Credit points:	2 CU x 1.59 = 3.18 ECTS
Prerequisites course(s):	-
Targeted learning outcomes:	CLO1 Understand the concepts of covalent bonds, ligands, stereochemistry, stability, magnetic properties and electronic spectra of coordination compounds. CLO2 Able to draw structures and predict the properties of coordination compounds CLO3 Communicate both verbally and in writing the concept of chemical bonds, stereochemistry, stability, magnetic properties, and electronic spectra of the coordination compound. CLO4 Demonstrate caring attitude and responsibility in applying coordination compounds in the environment
Content:	• The concept of coordination compounds, • Bond theories, • Types of ligands, • Stereochemistry, • Stability of complex ions, • Term symbols, • Multiplicity, • Orgel diagrams, • Tanabe-Sugano diagrams of coordination compounds

Study / exam achievements:	Students are considered to be competent and pass if at least get 55 Final score is calculated as follows: 20% participation + 30% assignment + 20% middle exam (UTS) & 30% final exam (UAS) Table index of graduation • A = 4 (85 ≤-≥ 100) • A- = 3,75 (80 ≤-< 85) • B+ = 3,5 (75 ≤- < 80) • B = 3 (70 ≤-< 75) • B- = 2,75 (65 ≤-<75) • C+ = 2,5 (60 ≤-<65) • C = 2 (55 ≤-<60) • D = 1 (40 ≤-<55) • E = 0 (0 ≤-<40)
Media:	Computer, LCD, White board
Learning Methods	Individuals assignment, group assignment, discussion, presentation
Literature:	1. Basolo, F and Johnson, R.C. 1986. Coordination Chemistry, 2nd Edition. New York: W.A. Benjamin, Inc. 2. Sugiarto, Bambang. 2006. Teori Senyawa Koordinasi. Surabaya: Unesa University Press. 3. Quagliano, J. V. And Vallarino, L. M., 1969. Coordination Chemistry, Massachusetts: D. C. Heath and Company 4. Huheey, E. James, Ellen, A.K, and Richard I.K. 1978. Inorganic Chemistry, Principle of Structure and Reactivity. USA: Harper Collins College Publishers 5. Madan, R.D., 1997. Modern Inorganic Chemistry , S. Chand and Company LTD, New Delhi.